

Stability of the Glycerol Facilitator in Detergent Solutions[†]

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ABSTRACT: Understanding membrane protein folding and stability is required for a molecular explanation of function and for the development of interventions in membrane protein folding diseases. Stable aqueous detergent solutions of the *Escherichia coli* glycerol facilitator in its native oligomeric state have been difficult to prepare as the protein readily unfolds and forms nonspecific aggregates. Here, we report a study of the structure and stability of the glycerol facilitator in several detergent solutions by Blue Native and sodium dodecyl sulfate (SDS)–polyacrylamide gel electrophoresis (PAGE), circular dichroism (CD), and fluorescence. Protein tetramers were prepared in neutral dodecyl maltoside (DDM) and in zwitterionic lysomyristoylphosphatidylcholine (LMPC) detergent solutions that are stable during SDS–PAGE. Thermal unfolding experiments show that the protein is more stable in LMPC than in DDM. Tertiary structure unfolds before quaternary and some secondary structure in LMPC, whereas unfolding is more cooperative in DDM. The high stability of the protein in DDM is evident from the unfolding half-life of 8 days in 8 M urea, suggesting that hydrophobic interactions contribute to the stability. The protein unfolds readily in LMPC below pH 6, whereas the tetramer remains intact at pH 4 in DDM. At pH 4 in DDM, the protein is more sensitive than at neutral pH to unfolding by SDS and the effect is reversible. At pH 3 in DDM, the tetramer unfolds, losing its tertiary structure but retaining native helical structure which melts at significantly lower temperatures than in the native tetramer. The glycerol facilitator prepared in SDS is mainly monomeric and has 10% less α -helix than the native protein. CD suggests that it forms a condensed structure with non-native tertiary contacts highly similar to the state observed in LMPC at low pH. The implications of the results for in vitro and in vivo folding of the protein are discussed.

Membrane proteins (MP)¹ are important elements of a wide variety of cellular processes, including the transport of molecules and ions, the reception and transduction of cell signals, and the transmission of light, electrical, and chemical energy. Despite their importance, much less is known about the structures, dynamics, and folding of membrane proteins than about their water-soluble counterparts. Some of the impediments to membrane protein analysis are low levels of cellular expression sometimes coupled with cytotoxicity upon overexpression, poor solubility in water, and a tendency to aggregate irreversibly (1–4). In addition, some membrane proteins appear to resist crystallization because of intrinsic conformational flexibility that is integral to their function (2). For example, lactose permease would not crystallize until

mutants were found that reduced its intrinsic flexibility (5), and the KvAP K⁺ channel required complexation with monoclonal Fab fragments for crystallization (6). The flexibility and tendency to aggregate suggest a low thermodynamic stability; however, a survey of diffraction data for MP indicates that they are at least as tightly packed as water-soluble proteins (7, 8). Direct measurements of membrane protein thermodynamic stability are problematic because of irreversible aggregation of the unfolded proteins (9, 10). Reversible chemical denaturation of a β -barrel protein has been reported (11), but in general, helical bundle proteins have proven to be more difficult to refold. It has recently been pointed out that helical bundle membrane proteins retain much of their secondary structure in sodium dodecyl sulfate (SDS), limiting its value as a chemical denaturant (12).

Because of the tendency of MP to aggregate, prokaryotic and eukaryotic cells have evolved a complex machinery to chaperone their folding and insertion into membranes. The current view of helical bundle protein folding is that transmembrane (TM) helix formation occurs in the aqueous channel of the translocon and that most or all of the tertiary structure formation occurs following partitioning of the protein into the lipid bilayer (9, 12–15). While the structure of the archaeal Sec Y translocon (16) has shed light on the process of membrane insertion, features of the sequence-specific code for transmembrane helices suggest that direct interactions with the membrane lipid are critical for insertion by the translocon (17). That the lipid bilayer plays a more

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¹ Abbreviations: AIDA, adhesin involved in diffusion adherence; AQP, aquaporin; Bis-Tris, bis(2-hydroxyethyl)aminotris(hydroxymethyl)methane; BN, Blue Native; CD, circular dichroism; DTT, dithiothreitol; DDM, dodecyl β -D-maltoside; EDTA, ethylenediaminetetraacetic acid; GF, glycerol facilitator; LMPC, lysomyristoylphosphatidylcholine; MALDI, matrix-assisted laser desorption/ionization; MP, membrane proteins; M_r , relative mass; MRE, mean residue ellipticity; NMR, nuclear magnetic resonance; NTA, nitrilotriacetic acid; OG, octyl β -D-glucopyranoside; PAGE, polyacrylamide gel electrophoresis; pI, isoelectric point; SDS, sodium dodecyl sulfate; TM, transmembrane; TOFMS, time-of-flight mass spectrometry; Tris, tris(hydroxymethyl)aminomethane; UV, ultraviolet.

than passive structural role in directing MP folding is also suggested by the fact that phosphatidylethanolamine is essential for the correct folding of lactose permease (18) and by molecular dynamics simulations (19). A two-stage working model of in vitro helix bundle MP folding was first proposed by Popot and Engelman (20–22) and then was expanded by White and Wimley (23). The broad outlines of the in vitro and in vivo models are surprisingly alike. In vitro folding begins with the partitioning of the protein to the membrane interface and is followed by α -helix formation, insertion of helices across the bilayer, and finally condensation to the native helix bundle (23). Quaternary structure might form through the association of monomers that have native tertiary structure or by the association of molten globule-like monomers that achieve their tertiary structure only in the oligomer (22).

The glycerol facilitator (GF) integral membrane protein from *Escherichia coli* is a member of the aquaporin family of proteins (24) and is an attractive model for studying membrane protein folding (25–27). GF permits the passage of glycerol and other low-molecular weight polyols, but not sugars, across the inner membrane of the bacterium (28–32). X-ray diffraction reveals a tetrameric, six-transmembrane α -helical protein that contains two additional half-membrane helices that line the glycerol pore (26). Although in vivo protein oligomer structure is often retained in crystals, determination of the structure of native oligomers is not straightforward even for water-soluble proteins (33, 34). The paucity of membrane protein structures has precluded systematic investigations of their quaternary structures, but it is worth noting that bacteriorhodopsin crystallizes both as a monomer (35) and as a trimer (36, 37). The oligomeric state of the GF has been the subject of numerous investigations by sedimentation on sucrose density gradients (38–40), cryo-electron microscopy (41, 42), freeze-fracture electron microscopy (43, 44), chemical cross-linking (27), and sodium dodecyl sulfate–polyacrylamide gel electrophoresis (SDS–PAGE) (27, 38–44). Most investigators have found the protein to exist as a monomer in SDS and as a mixture of oligomers in octyl glucoside, *N*-lauryl sarcosine, Triton X-100, and yeast, *E. coli*, and *Xenopus* membranes. This contrasts with the behavior of several of the homologous aquaporins that retain tetrameric structure even under the denaturing conditions of SDS–PAGE, including AQP0 (45, 46), AQP1 (47, 48), AQP2 (49), AQP4 (48), AQPcic (50), and AQP-Z (51). The generally accepted conclusion has been that GF is a weakly associating protein both in membranes and upon dissolution with detergents. However, it should be noted that a mixture of GF oligomers can also be prepared by adding urea to detergent-solubilized protein (52) and by sample aging, suggesting that the observed oligomers represent nonspecific associations of denatured protein induced by the extraction and/or measurement conditions. Finding conditions under which membrane proteins retain their native oligomeric state is important for dynamics studies, for understanding membrane protein folding, and in functional studies, as the oligomeric state of many membrane proteins is an important aspect of their regulation (53). Furthermore, although the seven-TM bacteriorhodopsin and several β -barrel porins have been thoroughly characterized, there have been no studies of the folding and stability of members of the aquaporin family (12).

We previously described the overexpression and characterization of the secondary structure of the glycerol facilitator in several detergents (27). Here, we describe for the first time two detergent preparations of the glycerol facilitator that show the same quaternary structure and stability as many of the aquaporins on SDS–PAGE. The native tetramer and low-pH states are characterized by denaturing and Blue Native (BN) polyacrylamide gel electrophoresis (PAGE), far- and near-ultraviolet circular dichroism (CD) spectropolarimetry, fluorescence spectroscopy, and sensitivity to pH, heat, urea, and SDS.

MATERIALS

Dodecyl β -D-maltoside was purchased from Anatrace (Maumee, OH) or Sigma (St. Louis, MO). Sodium dodecyl sulfate, octyl β -D-glucopyranoside, Coomassie Brilliant Blue R-250, β -mercaptoethanol, and urea were obtained from Sigma. 1-Myristoyl-2-hydroxy-*sn*-glycero-3-phosphocholine (LMPC) was from Avanti Polar Lipids (Alabaster, AL). Bis-Tris, Coomassie Brilliant Blue G-250, glycerol, and Tris were from Fisher Scientific (Fairlawn, NJ). Bromophenol blue was from BDH Inc. (Toronto, ON). Dithiothreitol was from Aldrich Chemical Co. (Milwaukee, WI), and imidazole was from Fluka (Buchs, Switzerland). Nickel-nitrilotriacetic acid (NTA) resin was from Qiagen (Toronto, ON). SDS–PAGE molecular weight markers were from Fermentas Life Sciences (Burlington, ON), and bovine serum albumin, the molecular weight standard for BN-PAGE, was from Sigma. All other materials were of the highest purity commercially available.

METHODS

Glycerol Facilitator Expression and Purification. The *E. coli* glycerol facilitator was expressed in BL21(DE3)pLysS cells (54) or C43(DE3) cells (55) from a pET28b(+) plasmid (Novagen) encoding an N-terminal His₆ purification tag and T7 epitope, as described previously (27, 30, 52). The M_r of the GF, including the N-terminal fusion tags, is 33505 Da and was confirmed by mass spectrometry (27); its calculated pI is 7.2 (MacVector, Oxford Molecular). For the purposes of purification of the glycerol facilitator, *E. coli* cells were mechanically sheared by brief sonication on ice, unbroken cells were removed by centrifugation, and membranes were prepared by ultracentrifugation at 100000g for 60 min, at 4 °C, in a Beckman sw28 rotor and either a Beckman L8 70 M or Beckman-Coulter LE 80k centrifuge (56). The membranes were solubilized in 25 mM sodium phosphate buffer (pH 7.6) containing 200 mM NaCl, 2 mM β -mercaptoethanol, and dodecyl β -D-maltoside (30 mM), octyl β -D-glucopyranoside (50 mM), sodium dodecyl sulfate (150 mM), or lysomyristoylphosphatidylcholine (20 mM), and the protein was purified using immobilized metal chelate chromatography. Protein purification was conducted by washing the resin with 15 mL of 25 mM sodium phosphate buffer (pH 7.5) containing 200 mM NaCl, 30 mM imidazole, and detergent (3 mM DDM, 25 mM OG, 2 mM LMPC, or 150 mM SDS). The resin was next washed with the same buffer containing 50 mM imidazole until the A_{280} was below 0.01. GF was eluted from the Ni²⁺-NTA column using the same buffer containing 250 mM imidazole. In the case of SDS-solubilized protein, the resin was washed with 5 mM

imidazole and the protein eluted using 50 mM imidazole or elution was conducted by lowering the pH as described previously (27). For the far-UV CD experiments, GF samples were dialyzed twice against buffer containing no imidazole. Adjustment of the pH was done by adding glacial acetic acid (up to 50 mM) followed by concentrated HCl, if necessary. Protein purity was analyzed by SDS-PAGE (57) and MALDI mass spectrometry (27).

Electrophoresis. For the SDS denaturing gels, protein was incubated in 50 mM Tris-HCl (pH 6.8), 2% SDS, 0.1% bromophenol blue, 10% glycerol, and 1% (v/v) β -mercaptoethanol for 10 min at 25 °C, prior to being loaded on the gel. Separation was achieved by SDS-PAGE in Laemmli discontinuous gels (57) composed of a 4% polyacrylamide stacking gel and a 10% resolving gel. After electrophoresis, proteins were visualized by staining with Coomassie Brilliant Blue R250.

For the Blue Native gels (58), protein was incubated in 80 mM Bis-Tris-HCl (pH 7.0), 400 mM 6-aminocaproic acid, and Coomassie Brilliant Blue G-250 (1/100 of detergent) for 10 min at 25 °C prior to being loaded on the gel. BN-PAGE was carried out using step gradient gels with a 4 to 10% polyacrylamide step gradient at 5 °C. Electrophoresis was started at 50 V and continued until all the protein entered the gel, after which the voltage was increased to 200 V for 3–4 h. Coomassie-stained polyacrylamide gel bands were quantified with *Un-Scan-It* digitizing software from Silk Scientific, Inc.

Circular Dichroism and Fluorescence. CD and fluorescence spectra were acquired with a Jasco 810 spectropolarimeter/fluorometer or a Jasco J500 A spectropolarimeter. The sample temperature was controlled by a Julabo F25 circulating water bath or with a built-in Peltier device. Briefly, for acquisition of far-UV CD spectra, protein solutions were placed in a quartz cuvette with a 0.1 cm path length and spectra were collected at 20 nm/min between 250 and 185 nm with a response time of 8 s and a data pitch of 0.1 nm. For near-UV spectra, the samples were placed in a quartz cuvette with a 1.0 cm path length and spectra were collected at 20 nm/min or 5 nm/min between 340 and 250 nm with a response time of 8 s or 30 s and a data pitch of 0.1 nm. Baselines were collected in the same fashion on buffer solutions, and spectra were baseline corrected. Protein concentrations were determined using the calculated molar absorptivity of 38305 M⁻¹ cm⁻¹ and are indicated in the figure legends. Note that due to the weak molar absorptivity of the aromatic amino acids, the protein concentrations required for near-UV CD measurements are on the order of 10-fold greater than those required for far-UV measurements. The CD intensity and wavelength of the spectropolarimeter were calibrated using solutions of *d*-10-camphorsulfonic acid (59). Mean residue ellipticities ($[\theta] \times 10^{-3}$ deg cm² dmol⁻¹) were calculated using the equation $[\theta]_M = M\theta/(10)(l)(c)(n)$, where *M* is 33505 g/mol, θ is the measured ellipticity in millidegrees, *l* is the cell path length, *c* is the protein concentration in grams per liter, and *n* = 316 residues. Deconvolution of the CD spectra into pure component spectra was performed using the algorithm CDSSTR (60) accessed through Dichroweb (61). For calculations of helix content, we have assumed that none of the 35 amino acids in the His₆-T7 tag are helical. Fluorescence spectra were collected in a 1 cm² rectangular quartz cell with a 0.5 s response time,

a data pitch of 1 nm, a scanning speed of 100 nm/min, and excitation and emission bandwidths of 8 and 10 nm, respectively. Thermal denaturation of membrane proteins is usually an irreversible process because the unfolded proteins aggregate and is therefore carried out for comparative purposes only. The thermal denaturation rate was 60 °C/h so that experiments using different probes and protein in different detergents would be self-consistent and comparable to experiments in the literature (62).

Curve Fitting. Near- and far-UV CD temperature titrations were fitted to the following equation (62, 63)

$$Y_{\text{obs}} = \frac{Y_f + m_f T + (Y_u + m_u T)e^{(-\Delta H_{\text{vH}}/R)(1/T - 1/T_m)}}{1 + e^{(-\Delta H_{\text{vH}}/R)(1/T - 1/T_m)}}$$

where Y_{obs} is the observed ellipticity, Y_f and Y_u are the ellipticities of the low- and high-temperature forms of the protein, respectively, m_f and m_u are the slopes of the curves at low and high temperatures, respectively, *R* is the gas constant, *T* is the temperature, and ΔH_{vH} is the van't Hoff enthalpy change of unfolding (62). The equation fits the thermal dependence of the ellipticity to a two-state unfolding transition yielding a value for T_m , the midpoint of the transition where the concentrations of folded and unfolded protein are equal. As the unfolding reactions are, in most cases, irreversible, the fits were used only to extract T_m values as quantitative indicators of the thermal stability of the protein (9, 10). The cooperativity of the transitions was quantified by calculating the range of temperatures (ΔT) over which 80% of the unfolding occurs and which corresponds to a K_{eq} range of 1/9–9. The urea unfolding rate constant was obtained by fitting changes in near-UV CD ellipticity at 268 nm to the following equation:

$$|Y_{\text{obs}}| = Ae^{-kt}$$

All the fits were done using the nonlinear least-squares fitting program in *Mathematica*, version 5.1 (64).

RESULTS

Glycerol Facilitator in DDM and LMPC at Neutral pH. In all previous reports of detergent solutions of the glycerol facilitator, the protein appeared predominantly monomeric by several assays, although SDS-PAGE sometimes showed varying amounts of dimer, trimer, tetramer, and higher-molecular weight species (27, 38–44). These observations have led to the suggestion that, in contrast to the aquaporins, the glycerol facilitators exist as monomers or weakly associating tetramers in membranes. In contrast, the results in Figure 1 show that recombinant GF, prepared by elution with imidazole from an immobilized Ni²⁺ resin at pH 7.0, exists predominantly as a tetramer in both DDM and LMPC according to SDS-PAGE (Figure 1A, lane 1 for DDM, and Figure 1C, lane 4 for LMPC). Because a two-Mg²⁺ ion coordination site was identified at the 4-fold symmetry axis of the GF tetramer crystal structure (26), we also investigated the separate effects of 100 mM EDTA, 50 mM Mg²⁺, and 50 mM Ca²⁺ on the tetramer. SDS-PAGE gave results identical to those shown in Figure 1A, ruling out the possibility that the tetramer is held together by metal chelation and further eliminating the possibility that the tetramer is an artifact of metal ion-mediated His₆ tag

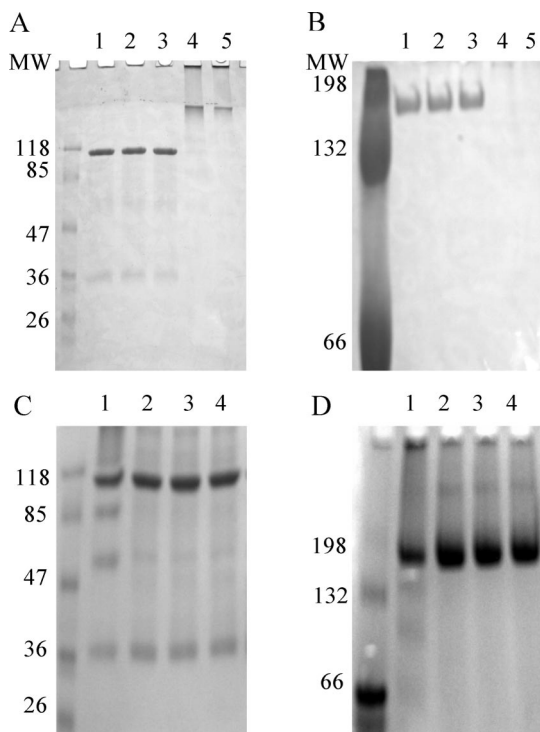


FIGURE 1: Thermal unfolding of GF quaternary structure by (A) SDS-PAGE and (B) BN-PAGE for DDM-solubilized protein (lane 1, 20 °C; lane 2, 40 °C; lane 3, 60 °C; lane 4, 80 °C; lane 5, 95 °C) and (C) SDS-PAGE and (D) BN-PAGE for LMPC-solubilized protein (lane 1, 95 °C; lane 2, 80 °C; lane 3, 60 °C; lane 4, 40 °C; lane 5, 20 °C). BN-PAGE MW standard: BSA dissolved in 3 mM DDM. The M_r of BSA is 66430 Da. GF tetramers electrophorese faster on SDS-PAGE than an unfolded water-soluble protein of the same size presumably because a folded membrane protein binds more SDS per residue than the unfolded water-soluble proteins.

association. The observations of tetramer via SDS-PAGE are also confirmed by electrophoresis of the same preparations on gentler Blue Native polyacrylamide gels (Figure 1B, lane 1 for DDM, and Figure 1D, lane 4 for LMPC). Since monomer, dimer, and trimer are virtually undetectable on the BN gels whereas the SDS gels reveal small amounts of the lower- M_r species, this suggests that the SDS present during SDS-PAGE causes slight dissociation of the tetramer. Interestingly, a small amount of an apparent octameric protein is sometimes observed in protein prepared both in DDM and in LMPC (see Figure 1D, lane 4).

CD and fluorescence spectroscopy were used to characterize the secondary and tertiary structures of the protein prepared at neutral pH in DDM and in LMPC detergent solutions. The far-UV CD spectra of GF dissolved in DDM at various temperatures, shown in Figure 2A, display the characteristics of a predominantly α -helical protein with negative bands at 219–222 and 209–211 nm, a positive band at 191–192 nm, and a crossover point at 200–201 nm. The ideal α -helical values are 222, 208, 190–195, and 205 nm, respectively (65). Deconvolution of the CD spectrum measured at 20 °C using the CDSSTR algorithm (60) yields 53% α -helix, 17% β -strand, 15% turns, and 15% unordered structure. Nearly identical values are obtained for GF dissolved in LMPC (Figure S1), where spectral deconvolution yields 53% α -helix, 10% β -strand, 15% turns, and 22% unordered structure. On the assumption that the 35-residue His₆-T7 tag is disordered, the deconvolutions indicate that 60% of the GF residues are helical in these detergents, a

result that is in close agreement with the secondary structure observed by X-ray diffraction where 61% of the residues are α -helical, 3% are 3_{10} -helical, and the balance are in turns or irregular structure (26).

Near-UV CD spectra of the glycerol facilitator tetramer solubilized in DDM at neutral pH and various temperatures are shown in Figure 2B. Near-UV circular dichroism is usually interpreted as indicating the presence of tertiary structure in proteins as only aromatic residues held rigidly in an asymmetric environment will display near-UV CD bands (66). The spectrum of GF in DDM shows strong, well-defined, negative bands between 255 and 290 nm and a positive band centered at 295 nm. Phenylalanine would be expected to contribute the least to the CD spectrum due to its high structural symmetry and the low sensitivity of its transitions to changes in solvent polarizability (67). Nevertheless, because of their relatively high abundance, the 21 phenylalanine residues in GF are the likely origin of the band (1L_b) at approximately 258 nm (Figure 2B). The small shoulder at 295 nm is most likely attributable to the 1L_b transitions of the seven tyrosine residues (65). Tryptophan residues usually dominate the near-UV CD spectra of proteins because of their large molar absorptivity, and the bands between 265 and 290 nm are attributable to the five tryptophan residues; the Trp 1L_a band is usually near 275 nm, and the 1L_b doublet usually falls between 280 and 290 nm (66). The near-UV CD spectrum of the protein dissolved in LMPC (Figure 2E) is nearly identical to that observed in DDM except that the weak band at 290 nm in DDM is positive and the positive band at 295 nm intensifies when the protein is dissolved in LMPC. The observation of well-defined CD absorbance bands suggests that the aromatic amino acids are fixed firmly in the structure in asymmetric environments and is a strong indication of a stable tertiary fold in both detergents.

Fluorescence spectra of GF dissolved in DDM are virtually identical to those measured in LMPC (see panels A and B of Figure S2). They both exhibit a peak at 329 nm and no evidence of Tyr emission at 303 nm. In folded globular proteins, tyrosine emission is often highly quenched by radiationless energy transfer to tryptophan as well as by interactions with other protein groups (68). The emission maximum for tryptophan in proteins is sensitive to both the polarity and the dynamics of the environment surrounding the side chain and is blue-shifted from 350 nm in environments of low polarity such as the hydrophobic interior of a protein or in a detergent micelle. Burstein and colleagues have classified Trp fluorescence in proteins into five classes depending on the polarity and dynamics of the Trp environments (69). According to their scheme, the dominant fluorescence at 329 nm in GF corresponds closely to a class of Trp side chains that are in a relatively nonpolar environment and H-bonded in a 2:1 exciplex that fluoresces at 331 nm (69). Inspection of the three-dimensional structure of GF (26) indicates that Trp-42 is well buried in the intermonomer interface and Trp-219 is buried but close to the surface of the protein. Two Trp side chains (76 and 215) project from the lateral surfaces of the protein and likely contact detergent but could also interact with water. In the crystal structure, Trp-48 forms part of the lining of the glycerol channel (26). However, it too is in a hydrophobic environment contacting Leu, Ile, Phe, and Val side chains and exposing only one

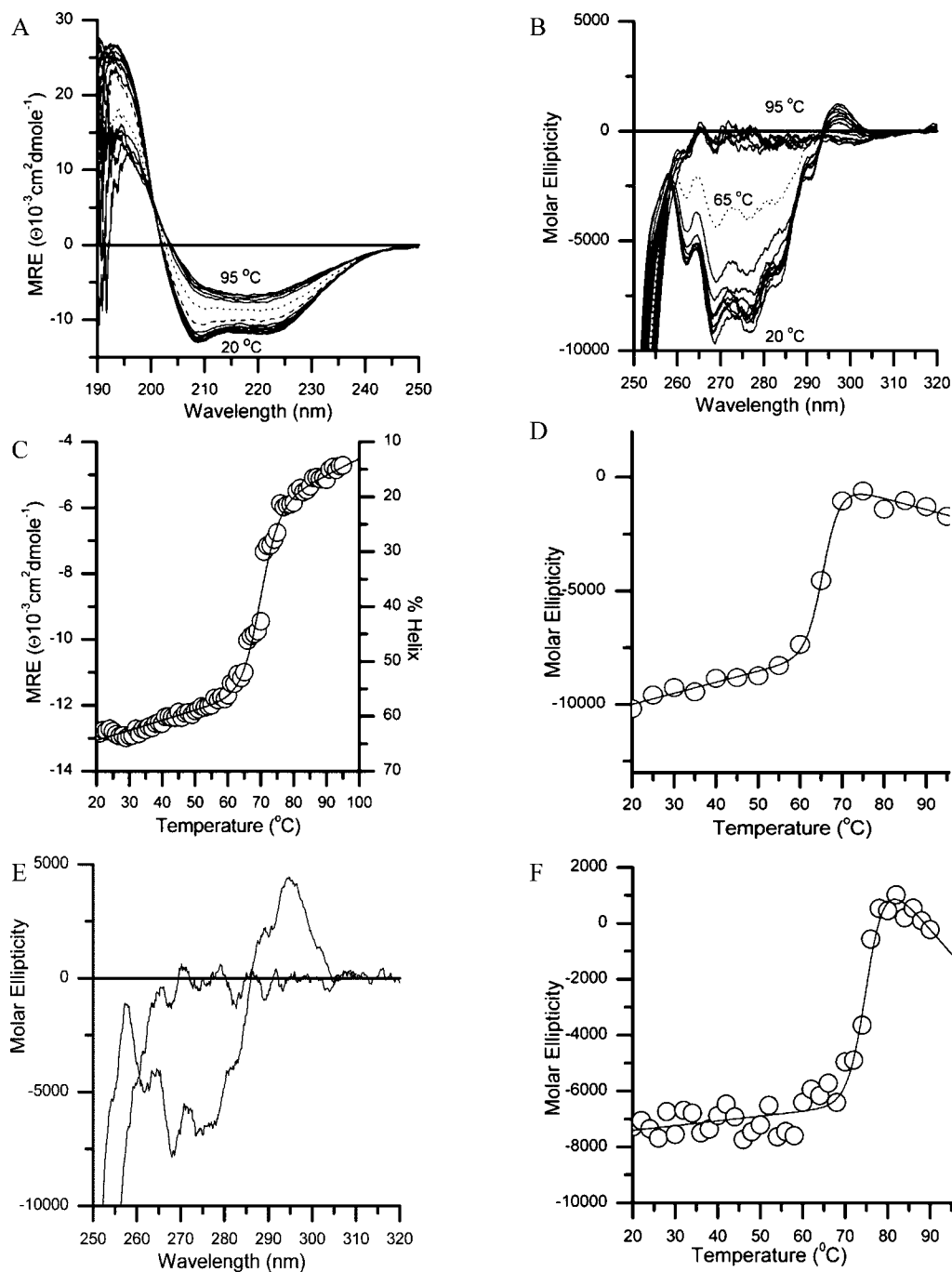


FIGURE 2: Thermal unfolding of secondary and tertiary structure of DDM- and LMPC-solubilized GF. (A) Far-UV CD spectra of 3.3 μ M GF in DDM between 20 and 95 $^{\circ}$ C in 5 $^{\circ}$ C intervals. (B) Near-UV CD spectra of 69 μ M GF in DDM between 20 and 95 $^{\circ}$ C in 5 $^{\circ}$ C intervals. (C) MRE at 209 nm from panel A as a function of temperature fit to a two-state unfolding transition as described in Methods. The fractional helicity was determined by deconvolution as described in Methods, and the results agree with the scale on the right side of panel C within $\pm 5\%$. (D) Molar ellipticity at 268 nm from panel B as a function of temperature fit to a two-state unfolding transition. (E) Near-UV CD spectrum of 25 μ M GF in LMPC at 20 $^{\circ}$ C and 95 $^{\circ}$ C. (F) Molar ellipticity at 268 nm from panel E as a function of temperature fit to a two-state unfolding transition.

edge of the ring to channel water. In summary, the fluorescence results are compatible with the observations of native secondary, tertiary, and quaternary structure described above and with the three-dimensional structure determined by X-ray diffraction (26).

Effects of Heat and Urea. To measure the stability of GF prepared in DDM and LMPC at pH 7, the effects of temperature and urea were examined. The thermal stabilities of the quaternary structure of the protein dissolved in DDM and LMPC were measured by SDS-PAGE (Figure 1A,C) and BN-PAGE (Figure 1B,D), and in each case, the protein

was heated at a rate of 1 $^{\circ}$ C/min. SDS-PAGE indicates that the tetramer in DDM is stable up to 60 $^{\circ}$ C. Above 60 $^{\circ}$ C, the protein unfolds, forming at least two high-molecular weight complexes; one aggregate enters both the stacking and running gels, and the other just enters the stacking gel (Figure 1A). The aggregation also eliminates the traces of monomer, dimer, and trimer that were present initially. Quantification of the gels indicates that at 60 $^{\circ}$ C only 9% of the tetramer has unfolded whereas at 80 $^{\circ}$ C 95% has unfolded. The results of BN-PAGE of the protein heated in DDM are shown in Figure 1B. The unfolding and aggrega-

tion of GF above 60 °C are evidenced only by the disappearance of the tetrameric protein from the gels, suggesting that the high- M_r oligomers are too large to enter the cross-linked polyacrylamide. This further suggests that the oligomers observed on the SDS gel in Figure 1A underestimate the actual oligomer size formed at high temperatures in DDM because of the ability of SDS to dissociate them. The thermal stability of the glycerol facilitator tetramer was also measured in LMPC. Both SDS-PAGE (Figure 1C) and BN-PAGE (Figure 1D) indicate that the GF tetramer is more thermally stable dissolved in LMPC than in DDM because 60% of the tetramer is still intact even at 80 °C.

The thermal unfolding of the protein's secondary structure in DDM was followed by far-UV CD spectropolarimetry, and the spectra are shown in Figure 2A. The isodichroic point at 201 nm suggests that a two-state transition is occurring. Plots of the temperature dependence of the ellipticity at 209 nm fit to a two-state unfolding transition are shown in Figure 2C. Although the unfolding transitions are not reversible, the transition midpoints can be used to quantify the thermal stability of the protein (9). The spectra indicate that in DDM the protein secondary structure is stable up to 60 °C, after which increasing temperature reduces the fractional helicity from 60 to 15% with a T_m of 69.7 ± 0.7 °C and a transition width, ΔT , of 9.8 °C. The near-UV CD ellipticity is also relatively unperturbed up to 60 °C (Figure 2B,D) and is eliminated with a T_m (65.5 ± 0.7 °C) and a ΔT (7.5 °C) that are only slightly lower than those measured by far-UV CD. The heating rates for the far- and near-UV CD experiments were 0.43 and 0.28 °C/min, respectively. Increasing the rate of heating to 1 °C/min increased the melting points of the far- and near-UV CD transitions to 75 ± 1.1 and 71 ± 2 °C, respectively. This result is in reasonably good agreement with the observations of Sehgal et al. (62), who measured the dependence of T_m on the rate of heating in the *E. coli* AIDA (adhesin involved in diffusion adherence) protein. They reported that the T_m of AIDA is elevated by 6 °C when the scan rate is increased from 0.25 to 1 °C/min but that the T_m is more sensitive to heating rates below 0.66 °C/min and less sensitive at rates above this.

The thermal stability of the tertiary structure was also monitored by Trp fluorescence; elevated temperatures progressively quench the Trp fluorescence, and a small red shift from 329 to 333 nm is detected in the spectra at temperatures above 60 °C (Figure S2A). Thus, as the tertiary structure unfolds, there is a slight increase in accessibility to water of one or more of the Trp residues. Taken together with the electrophoresis data (Figure 1A,B), the CD and fluorescence results suggest that the thermal transition in DDM involves a concomitant loss of secondary, tertiary, and quaternary structure with a transition midpoint between 66 and 70 °C.

Compared to the protein dissolved in DDM, the glycerol facilitator secondary and tertiary structures in LMPC show significantly higher thermostability. When measured at a heating rate of 1 °C/min, the secondary structure unfolds with a T_m of 86.9 ± 11.5 °C; however, the transition width (ΔT) is wide (15.1 °C), suggesting helices with varying thermal stability (panels A and B of Figure S1). According to near-UV CD, the tertiary structure of the protein unfolds at 75.0 ± 1.5 °C with a transition width (ΔT) of 9.2 °C in LMPC (Figure 2F). These measurements were made by heating the samples at 1 °C/min,

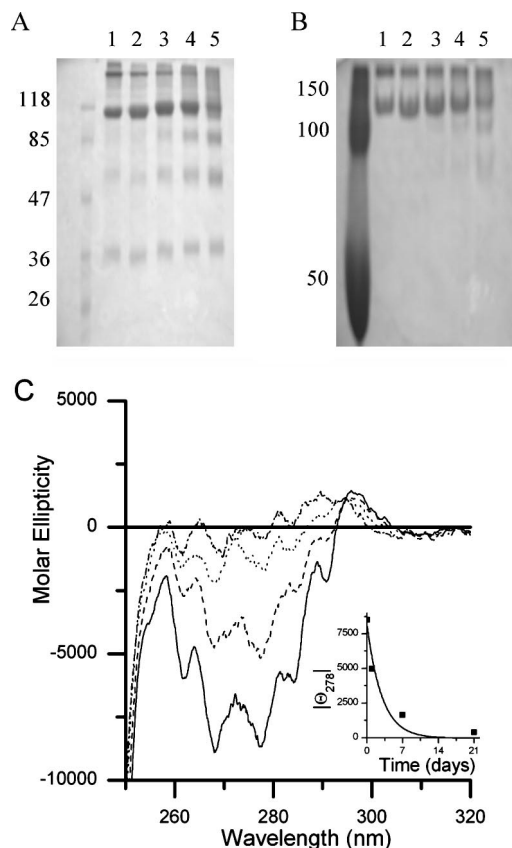


FIGURE 3: Effects of urea on the quaternary and tertiary structure of DDM-solubilized GF at pH 7. Protein was incubated in urea for 3 weeks and then analyzed by (A) SDS-PAGE and (B) BN-PAGE: lane 1, 0 M urea; lane 2, 2 M urea; lane 3, 4 M urea; lane 4, 6 M urea; lane 5, 8 M urea. (C) Time dependence of unfolding of 55.4 μ M GF tertiary structure by 8 M urea: (—) 1 h, (---) 24 h, (····) 7 days, and (-·-·) 21 days. The inset shows the fit of the change in ellipticity at 268 nm to an exponential decay as described in Methods.

the same rate used for the DDM experiments and the same rate used by Sehgal et al. (62). Measurements were also taken at 0.5 °C/min and resulted in a lowering of the near-UV CD T_m by 1.6 °C. Thus, the protein in LMPC appears slightly less sensitive to heating rate than in DDM. Overall, LMPC appears to stabilize the secondary structure more than the tertiary structure and lowers the cooperativity of the unfolding between the levels of structure. Interestingly, the quaternary structure appears to unfold concomitant with the secondary structure rather than the tertiary structure as the T_m is well above 80 °C.

Urea is a common protein denaturant used to measure the stability of many water-soluble proteins and has recently been shown to reversibly unfold a β -barrel protein (11). Initial experiments with DDM-solubilized GF indicated a time dependence of the effects of urea, and in an attempt to establish equilibrium conditions, the protein was incubated in various concentrations of urea for 3 weeks. Panels A and B of Figure 3 show the effects of additions of urea on the DDM-solubilized protein quaternary structure. According to SDS-PAGE, low concentrations of urea (2–4 M; lanes 2 and 3 in Figure 3A) have an only minor effect on the amount of tetramer present in the solution. Small increases in monomer, dimer, and trimer are observed, but the most notable change is the significant loss of octamer. After incubation for 3 weeks in 8 M urea (Figure 3A, lane 5), the tetramer is nearly absent, the levels of monomer, dimer, and trimer have increased substantially, and the amount

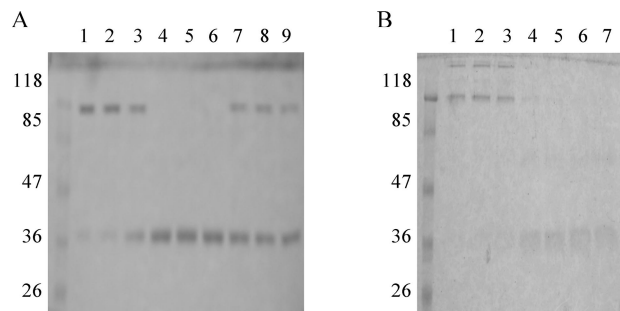


FIGURE 4: Effects of increasing and decreasing the pH (A) and incubation with SDS at pH 4 (B) on the quaternary structure of DDM-solubilized GF. (A) SDS-PAGE: lane 1, pH 7.8; lane 2, pH 7.0; lane 3, pH 6.0; lane 4, pH 5.0; lane 5, pH 4.0; lane 6, pH 5.0; lane 7, pH 6.0; lane 8, pH 7.0; lane 9, pH 7.8. (B) SDS-PAGE: lane 1, 0 mM SDS; lane 2, 2 mM SDS; lane 3, 5 mM SDS; lane 4, 10 mM SDS; lane 5, 20 mM SDS; lane 6, 40 mM SDS; lane 7, 60 mM SDS.

of octamer is greatly reduced. In addition, material electrophoresing between the tetramer and octamer appears that is likely an unfolded aggregate soluble in high concentrations of SDS and urea. BN-PAGE of the same samples is shown in Figure 3B. The BN gels are less sensitive to the unfolding induced by urea at low concentrations because in the absence of SDS the low-molecular weight species cannot be observed on the gels; presumably, they form aggregates too large to enter the gel. It is also likely that in the absence of SDS the tetramer is more resistant to dissociation in low concentrations of urea. However, at 6 and 8 M urea, the unfolding of the tetramer and octamer is evident from their diminishing bands on the BN gel and from the very faint bands that appear to be trimer and dimer that appear. The latter are soluble in small amounts because of the high concentrations of urea present, whereas in DDM alone, they are not observed.

Figure 3C shows another indication of the resistance of the protein to denaturation by urea. Incubation of the DDM-solubilized GF in 8 M urea for 1 day causes very little change in the near-UV CD spectrum. However, over the course of 21 days, the added urea reduces the level of tertiary structure with an exponential unfolding rate constant of $(4 \pm 1) \times 10^{-6} \text{ s}^{-1}$. The high concentrations of urea required to unfold the protein and the slow unfolding rate both suggest a high thermodynamic stability of the protein in DDM.

Influence of pH and SDS on GF Structure. The electropherogram in Figure 4A shows the influence of pH on the quaternary structure of DDM-solubilized GF initially prepared at pH 7.8. Lane 1 shows that the protein is predominantly tetrameric at pH 7.8 with a small amount of monomer present on SDS-PAGE. Lowering the pH to 7.0 (lane 2) causes little change in the oligomer structure, but at pH 6.0 (lane 3), the tetramer band is diminished and the monomer band increased, indicating partial dissociation of the tetramer into monomer. At pH 5.0, the tetramer is completely converted to monomer (Figure 4A, lane 4), and lowering the pH to 4.0 causes no further change in the oligomerization state (Figure 4A, lane 5). It is worth noting that the lanes containing monomer show no dimer and trimer protein, suggesting that the monomer is in an unaggregated and partially folded state. To determine if the acid-induced unfolding of the tetramer could be reversed, the pH of the same protein preparation was elevated and the refolding monitored by SDS-PAGE. When the pH is increased from

4.0 to 5.0 (lane 6), only monomeric DDM-solubilized protein is observed in accord with what was observed when the pH was lowered from pH 6.0 to 5.0 (lane 4). When the pH is increased from 5.0 to 6.0, 60% of the monomer is restored to the tetramer state (lane 7), indicating substantial reversibility of the pH-induced unfolding. Elevation of the pH to 7.0 (lane 8) and 7.8 (lane 9) causes little conversion of monomer to tetramer. It seems unlikely that the tetramer band is an unfolded protein aggregate as no dimer or trimer appears on the electropherogram. Note also that some of the protein that is irreversibly unfolded appears on the gels as a high- M_r aggregate near the top of the gel.

Whenever SDS-PAGE is used to monitor quaternary structure, the protein is subjected to approximately 60 mM SDS from the sample treatment buffer. In an attempt to separate the effects of pH and SDS on the pH-induced unfolding, an SDS titration was carried out. DDM-solubilized GF was incubated at pH 4.0 with 0–60 mM SDS for 10 min, and then the pH was increased to 7.0. In the absence of SDS (Figure 4B, lane 1), GF electrophoreses as a tetramer with very little monomer observed when the pH is returned to 7.0. Interestingly, cycling the pH induces the formation of small amounts of octameric protein. The presence of 2 mM SDS has no measurable effect on the oligomeric state of the protein (lane 2), whereas at 5 mM SDS (lane 3), monomer and dimer can be observed in the preparation. At 10 mM SDS (lane 4), the majority of tetrameric and octameric GF is lost and significant amounts of both monomer and dimer are observed. Further increases in SDS concentration to 20, 40, and 60 mM (Figure 3, lanes 5–7, respectively) continue to decrease the levels of the tetrameric and octameric species while increasing the levels of the monomeric and dimeric forms of the pH 4 DDM-solubilized GF. These experiments suggest that the SDS present in the SDS-PAGE sample buffer induced the tetramer dissociation observed in Figure 4A, and BN-PAGE confirmed that in the absence of SDS no monomer is formed at pH 4 (data not shown). These results show that there is a pH-dependent change in the structure of the protein making it more susceptible to SDS unfolding but that these changes are at least partially reversible.

The effects of pH on the structure of the GF tetramer dissolved in LMPC are markedly different compared to the effects on the protein dissolved in DDM. SDS-PAGE (Figure S3A) shows the pH-dependent replacement of the tetramer with monomer, dimer, and trimer in LMPC, indicative of irreversibly unfolded protein, whereas in DDM, only monomeric protein is observed at low pH. BN gels (not shown) confirm the sensitivity of the LMPC-solubilized protein to pH. At pH 4, the tetramer has almost completely disappeared from the gel, and in its place are low-molecular weight species (monomer, dimer, and trimer) and a very high-molecular weight species which enters the top of the stacking gel.

Changes in secondary and tertiary structures brought about by lowering the pH were investigated spectroscopically. In DDM and in LMPC, pH caused essentially no change in the secondary structure of the protein between pH 7.8 and 3 as judged by far-UV CD (panels A and B of Figure S4). Similarly, lowering the pH of the protein dissolved in DDM to pH 4 diminishes only slightly (20%) both the positive and negative ellipticity in the near-UV CD spectrum (Figure

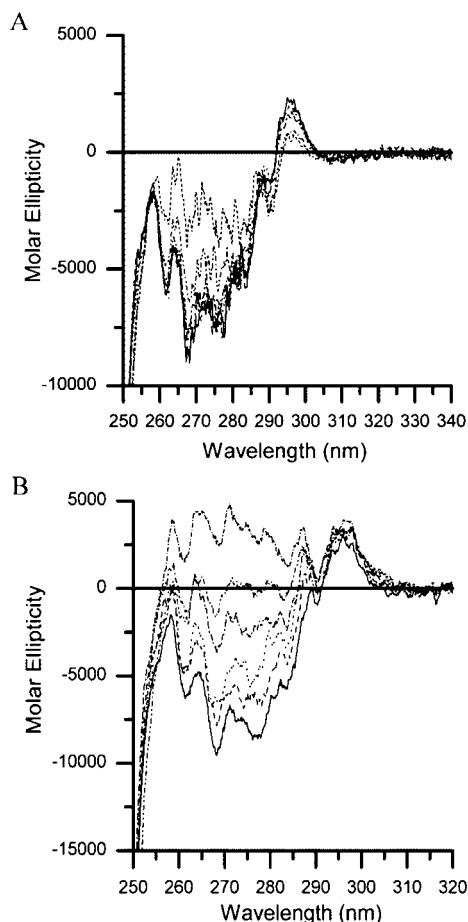


FIGURE 5: Effects of pH on the tertiary structure of DDM- and LMPC-solubilized GF: (—) pH 7.0, (---) pH 6.0, (···) pH 5.0, (-·-) pH 4.0, and (-··-) pH 3.0. (A) GF (69 μ M) in DDM, with tertiary structure monitored by near-UV CD. (B) GF (80 μ M) in LMPC, with tertiary structure monitored by near-UV CD.

5A). However, between pH 4 and 3, a more significant change occurs eliminating most of the near-UV CD signal intensity (Figure 5A). In contrast, in LMPC (Figure 5B), the near-UV CD band at 295 nm increases in intensity as the pH is lowered and the ellipticity between 255 and 290 nm changes from negative at pH 7 to positive at pH 3. Also notable is the significant change in the CD signal between pH 7 and 4 indicating a greater sensitivity to the acidic pH of the protein in LMPC.

The effects of pH on the fluorescence of the protein were also measured (panels A and B of Figure S5). In DDM, between pH 7 and 5, the Trp emission is quenched by 15% but there is no detectable change in the position of the emission maximum. However, between pH 5 and 4, the quantum yield is restored to slightly more than its original value and the peak maximum shifts 4 nm to the red. Between pH 4 and 3, the fluorescence is quenched by 30% but the emission maximum remains at 333 nm. In LMPC, decreasing the pH continuously quenches the Trp emission between pH 7 and 4; the emission maximum shifts to the red by 2 nm between pH 7 and 6 and shifts a further 2 nm between pH 6 and 5. There is little change in either the quantum yield or the emission maximum between pH 4 and 3. These changes likely reflect direct effects of the protonation of different side chains in the local environments of the five Trp residues as well as indirect effects of the unfolding of the tertiary structure as the pH is decreased.

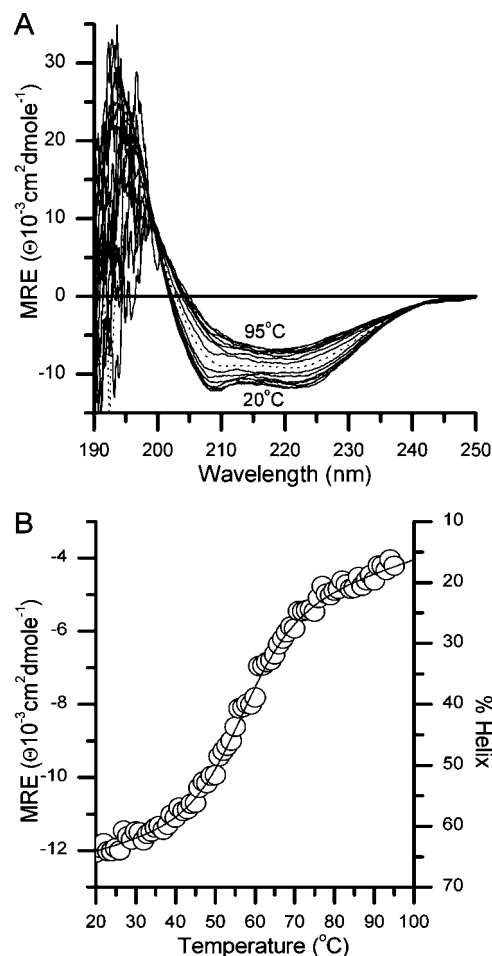


FIGURE 6: Characterization of the pH 3.0 DDM-solubilized GF molten globule by CD. (A) Thermal denaturation of secondary structure monitored by far-UV CD (2.6 μ M). Spectra were collected between 20 and 95 $^{\circ}$ C in 5 $^{\circ}$ C intervals. The dotted line is the 55 $^{\circ}$ C spectrum. (B) MRE at 209 nm as a function of temperature fit to a two-state unfolding transition as described in Methods. The fractional helicity was determined by deconvolution as described in Methods, and the results agree with the scale on the right side of Figure 6B within $\pm 5\%$.

At pH 3, the GF electrophoreses as a monomer via SDS-PAGE and as a mixture of monomer and high-molecular weight oligomers via BN-PAGE (not shown), retains its native content of α -helix (Figure S4A), but has significantly reduced tertiary structure content in DDM (Figure 5A). Mildly acidic conditions (pH 3–4) have been used to produce molten globule states of a large number of water-soluble proteins (70). The stability of the pH 3 state of the glycerol facilitator dissolved in DDM was studied by unfolding it with heat. In Figure 6A, the CD spectra show that the protein at low pH unfolds with a significantly lower T_m (56.9 ± 1.4 $^{\circ}$ C) and over a wider temperature range ($\Delta T = 23.4$ $^{\circ}$ C) than the neutral-pH form. Deconvolution of the CD spectra indicates that at 95 $^{\circ}$ C 18% of the core of the α -helix is still folded in the pH 3 state of the protein. The thermal unfolding was also monitored by fluorescence; as the temperature is elevated, there is a progressive quenching of Trp fluorescence but no change in the emission maximum at 333 nm (data not shown). SDS-PAGE shows that, similar to the case for the high-pH form of the protein, thermal unfolding of the low-pH form causes nonspecific aggregation of the protein into high-molecular weight oligomers (data not shown).

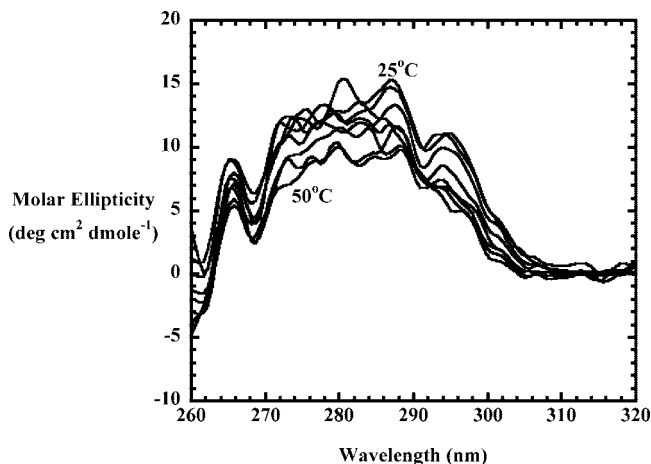


FIGURE 7: Effect of temperature on the near-UV CD spectra of GF dissolved in SDS. Protein was purified as described previously (27), and the concentration was 77 μ M. Spectra were recorded at 20 and 25 $^{\circ}$ C and between 30 and 80 $^{\circ}$ C in 10 $^{\circ}$ C intervals.

Glycerol Facilitator in SDS. We have previously shown that GF prepared in SDS and eluted from the Ni^{2+} resin at pH 4.2 electrophoreses predominantly as a monomeric protein on SDS–PAGE and contains an α -helix content of 48% (53% if one accounts for the affinity tag) (27). The α -helix in low-pH-eluted, SDS-solubilized protein is thermally stable up to 50 $^{\circ}$ C, and at 70 $^{\circ}$ C the protein still retains 40% of its helix content (data not shown). Thus, in SDS the α -helices are thermally very stable which has been noted for several other membrane proteins (12). Near-UV CD spectra of GF dissolved in SDS at various temperatures are shown in Figure 7. The spectra are similar to that observed for the protein at pH 3 in LMPC (Figure 5B). The weak positive ellipticity and the less well-defined absorption bands suggest some transient asymmetry in the environments of the aromatic residues. Increasing the temperature does not eliminate the spectrum, and there is therefore no evidence of a two-state unfolding transition. The protein in SDS thus appears to reside in a compact association of α -helices retaining much of its native secondary structure, has an altered tertiary structure, but retains none of its quaternary structure.

DISCUSSION

GF Oligomers. In cells, membrane protein folding and the stability of the native state depend on a complex interplay among amino acid sequence, water, lipid hydrocarbon, lipid head groups, and the translocon. Despite the importance of environmental variables, it is unlikely that the native states of membrane proteins are kinetically trapped intermediates determined by the folding pathway. Rather, just as in the case of water-soluble proteins that fold with the aid of chaperones (71), membrane proteins likely reside at thermodynamic free energy minima dictated mainly by amino acid sequence. In vitro folding studies of membrane proteins can therefore provide information about the intra- and intermolecular forces that stabilize the native fold if it can be isolated (12). In contrast to earlier reports (27, 38–44), we have shown here that the glycerol facilitator exists as a stably folded tetramer in solutions of the neutral detergent DDM and in zwitterionic LMPC. That the tetramer is the native state of the protein is supported by observations of

tetramers in the crystalline state by X-ray diffraction (26) and by cryoelectron microscopic observations of two-dimensional crystals of protein reconstituted in *E. coli* lipid (42). There is some evidence to suggest that the glycerol facilitator exists in close association at the inner membrane (72) with the *E. coli* glycerol kinase, which is active both as a dimer and as a tetramer. A 1:1 association between the two protein tetramers would provide the most efficient pathway for transferring glycerol from the facilitator to the kinase, preventing loss from the cytoplasm of membrane-permeable glycerol (26, 73).

In addition to the tetramer, we also observe varying amounts of octameric glycerol facilitator in the detergent preparations. The octamers appear to be less stable than the tetramers as heat and urea dissociate the octamer under conditions milder than those used for the tetramer. Scanning-tunneling electron microscopy of freeze-dried solubilized GF showed that significant amounts of octamer coexist with tetramers and even a small fraction of dodecamer (42). Octamers and higher associations of aquaporin-0 have been observed by BN-PAGE, by analytical ultracentrifugation, and in loosely packed three-dimensional crystals (74). The structure of the aquaporin-0 octamer suggests that a plausible function is the formation of cell contacts between closely apposed lens cells. Weak self-association of GF tetramers in detergent preparations suggests the possibility that the GF protein may cluster in the inner membrane of *E. coli* and that the clustering may be functionally significant (75).

GF Thermal Stability. Most membrane proteins aggregate irreversibly upon thermal unfolding, precluding a rigorous thermodynamic treatment. However, because unfolding and oligomerization are often coupled, the onset of oligomerization can give an indication of thermal stability (9) and may yield valuable information about the folding pathway. A recent review of membrane protein stability indicates that helical membrane proteins unfold with low enthalpies of denaturation compared to water-soluble proteins, that transition midpoints depend on the nature of the solubilizing detergent or lipid, and that a core of highly stable α -helical structure often remains at elevated temperatures (9, 76). For example, bacteriorhodopsin unfolds with a T_m of 100 $^{\circ}$ C in purple membranes, leaving much of its helix content intact, whereas the T_m is reduced to 80 $^{\circ}$ C in detergent (for a review, see ref 76). Cytochrome oxidase unfolds in dodecyl maltoside in two stages with T_m values of 47 and 67 $^{\circ}$ C, leaving 70% of its helix content intact at high temperatures (77). Similar results with other proteins suggest that the surface-exposed loops and domains of helical membrane proteins unfold at lower temperatures, producing a molten globule-like structure of loosely associated, thermal stable α -helices. In general, such studies support the models of in vitro (20–23) and in vivo (9, 10, 14–16) MP folding in which tertiary structure forms by an association of preformed α -helices.

In this work, far-UV circular dichroism (Figure 2A) shows that the helix contents of the glycerol facilitator measured in DDM and in LMPC solutions at neutral pH and 25 $^{\circ}$ C are identical and in close agreement with the structure observed in the crystals (26); fluorescence (panels A and B of Figure S2) and near-UV circular dichroism (Figure 2B,E) show that the tertiary structures are nearly identical in the two detergent solutions. However, significant differences are observed in the sensitivity of the protein to heat and acid in

the two detergents. In DDM, the protein appears to unfold secondary, tertiary, and quaternary structures in a more or less cooperative transition between 71 and 75 °C. In LMPC, all levels of structure are more thermally stable than in DDM and the unfolding is less cooperative. The melting points of the secondary structure ($T_m = 87$ °C) and quaternary structure ($T_m > 80$ °C) are similar and significantly higher than that of the tertiary structure ($T_m = 75$ °C); however, the secondary structure unfolding transition ($\Delta T = 15$ °C) is broad, suggesting helices with varying stability. The lower stability of the protein in DDM is also indicated by the lower thermally stable helix content remaining at 95 °C in DDM (14%) compared to LMPC (22%). The loss of tertiary structure before secondary structure in LMPC and the presence of a core of thermally stable helices in both detergents support the standard view that the connecting loops are more thermally labile than the α -helices. However, the cooperative unfolding of much of the helix content along with the tertiary and quaternary structure in DDM indicates the influence that detergent can have on the unfolding pathway. Interestingly, high concentrations of urea unfold the protein dissolved in DDM only very slowly (Figure 3). It has been argued that urea does not disrupt hydrophobic interactions (78), and this suggests that hydrophobic interactions between the protein monomers play a major role in stabilizing the protein structure.

A possible explanation of the higher thermal stability of the GF in LMPC compared to DDM is that the phosphocholine headgroup provides an electrostatic environment that is better matched to GF at pH 7 than the neutral maltoside. This seems likely in light of the high abundance (>80 mol %) of phosphatidylethanolamine in the *E. coli* inner membrane which is iso-Coulombic with phosphatidylcholine. In LMPC, the quaternary structure of the GF appears to be more thermally stable than the tertiary structure. This suggests that, in vivo, the tetramer may form by the association of molten globule-like structures that contain native α -helices but no fixed tertiary structure and that compact tertiary folds form only in the assembled tetramer. This pathway would assign a chaperone function to each of the monomers and highlights the importance of the tetramer in stabilization of the protein tertiary structure. Such a mechanism may be widespread as many membrane proteins are oligomeric and may require multiple protein–protein contacts for formation and stabilization of well-packed tertiary cores. Dimeric pro-caspase-3, a water-soluble protein, has been shown to unfold to an associated dimer intermediate, and the stability of the dimer contributes a significant fraction of the conformational free energy of the protein (79).

The significant difference in stability between bacteriorhodopsin and the glycerol facilitator might be due to the structures of the proteins. The seven-helix bundle protein BR consists of seven α -helices surrounding a central retinal, and the bound hydrophobic prosthetic group likely contributes to the added stability. The GF also consists of a bundle of seven transmembrane α -helices; however, one helix is comprised of two half-helices that span only half the bilayer, returning to the surface of the protein as an irregularly folded loop (26). In the absence of compensating interactions, the two half-helices are expected to be less stable than a full TM helix because of a smaller fraction of central residues fixed by two helical H-bonds per residue and a reduced

number of stabilizing dipole–dipole interactions. However, this structural difference cannot explain the apparent greater stability of the aquaporins in solution (45–51) compared to the glycerol facilitator (27, 38–44) as the protein folds are very similar. A full explanation will require a more detailed study of the stability of the aquaporins.

GF pH Stability. The GF shows significantly different stability to acid pH in the two different detergents that were tested. In zwitterionic LMPC, the protein irreversibly unfolds and aggregates as the pH is lowered from 7 to 4 (Figure S3A). Although the charge state of the detergent phosphocholine is unchanged over this pH range, the buildup of positive charge on the protein must be incompatible with the charges on the detergent headgroup. In the neutral detergent DDM, all levels of structure appear to remain folded down to pH 4. However, at pH 4 in DDM, the tetramer protein is less stable than at pH 7 as small additions of SDS are enough to dissociate it to a monomeric state (Figure 4B). The conformational change must be subtle as it can be detected only by changes in Trp fluorescence (Figure S5A) and minor changes in the near-UV CD spectrum (Figure 3C). Furthermore, the structural changes that occur in the protein between pH 7 and 4 appear to be partially reversible, whereas decreasing the pH to 3 irreversibly unfolds the tetramer and protein tertiary structure. At pH 3 in DDM, the protein is monomeric, retains its native helix content, but has lost most of its tertiary organization. This “molten globule-like” association of helices has a lower thermal stability than the native tetramer at neutral pH but still retains 18% of its native helix core at 95 °C. Mildly acidic conditions (pH 3–4) have been used to produce molten globule states of a large number of water-soluble proteins (70), and the equilibrium molten globule is thought to be an analogue of an important kinetic intermediate in the folding pathway of many water-soluble proteins (70). It has also been shown to be an important intermediate on the membrane insertion pathway of the colicin Ia channel domain (80).

The reversible unfolding of several membrane proteins has been accomplished with the use of SDS, but recently, it has been pointed out that SDS has a strong helix stabilizing capacity and that chemical unfolding with SDS is unlikely to produce a disordered polypeptide (12). Instead, membrane proteins in SDS likely exist in condensed states with their tertiary structure disrupted by intercalation of SDS molecules between the TM helices. The near-UV CD spectrum of the GF dissolved in SDS shows mainly positive bands with significant band structure suggestive of the maintenance of some tertiary contacts. The spectrum in SDS is strikingly similar to that observed in LMPC at pH 3, suggesting that in LMPC at low pH the protein forms a condensed unfolded structure similar to that present in SDS.

These results show that the glycerol facilitator is less thermally stable than bacteriorhodopsin but more stable than cytochrome oxidase. In the case of the glycerol facilitator, the measured stability is in good agreement with the observations of a tightly packed three-dimensional structure and low-temperature *B*-factors by X-ray diffraction (26). The results also agree with calculations of root-mean-square deviations of C α atoms during molecular dynamics simulations that indicate only very minor conformational changes during the passage of glycerol through the pore (81, 82). The molecular dynamics simulations indicate that there is

no cooperativity between the subunits of the tetramer during glycerol passage (81), and this suggests that the tetramers play a structural role in stabilizing the monomer conformation in the hydrophobic membrane. The molecular dynamics simulations also point out that the stability of the pore conformation is critical both for the conduction of glycerol and for the selectivity of the pore for glycerol over water (81). Kinetically and thermodynamically stable preparations of membrane proteins in detergent may be amenable to structure and dynamics analysis by multinuclear NMR TROSY experiments (83) if sufficient deuterium-labeled protein can be prepared at concentrations of 200–500 μ M (84).

SUPPORTING INFORMATION AVAILABLE

Thermal unfolding of secondary structure LMPC-solubilized GF (Figure S1), thermal unfolding of GF monitored by tryptophan fluorescence (Figure S2), effects of pH on the quaternary structure of GF dissolved in LMPC (Figure S3), effects of pH on the far-UV CD spectra of GF (Figure S4), and effects of pH on the tryptophan fluorescence of GF (Figure S5). This material is available free of charge via the Internet at <http://pubs.acs.org>.

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